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solution in chemistry a homogenous mixture of two or more substances in relative amounts that can be varied continuously up to what is called the limit of solubility the term solution is commonly applied to the liquid state of matter but solutions of gases and solids are possible in chemistry a solution is a type of homogeneous mixture consisting of two or more substances in which one substance is dissolved in another the solvent because the mixture is homogeneous a sample of a solution has the same appearance and concentration as any other sample what is a solution a solution is a homogeneous mixture of two or more pure substances the substance that is in a large amount in the solution is called the solvent the substance that is in smaller amounts in a solution is called the solute in chemistry a solution is a special type of homogeneous mixture composed of two or more substances in such a mixture a solute is a substance dissolved in another substance known as a solvent let us take a closer look at what we mean by a solution starting with a two component system typically one of the components is present in a smaller amount than the other we call the major component the solvent and the minor component s the solute s the most familiar solutions are aqueous solutions in which water is the solvent salt dissolved in water is a solution the major component of a solution called the solvent is typically the same phase as the solution itself each minor component of a solution and there may be more than one is called the solute a solution is what occurs when two chemicals are mixed referred to as a solvent and a solute the combination of solute and solvent together is a solution these are often one substance dissolved in another such as dissolving salt in water although a solution can be made from any state of matter combining a solute and a solvent makes a solution solutions are used in a wide range of science experiments including chemistry biology health cooking and materials science this video and written tutorial will guide you through how to make solutions and how to calculate the concentration of a solution a solution is a homogeneous mixture of two or more substances a solution may exist in any phase a solution consists of a solute and a solvent the solute is the substance that is dissolved in the solvent the amount of solute that can be dissolved in a solvent is called its solubility a solution is a homogeneous mixture of two or more components in which the particle size is smaller than 1 nm common examples of solutions are sugar in water and salt in water solutions soda water etc in a solution all the components appear as a single phase sieving a mixture made of solid particles of different sizes for example sand and gravel can be separated by sieving separating a mixture of iron filings and sand a mixture of iron filings and sand

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~~solution is a homogeneous mixture of two or more substances the term~~  
homogeneous means the same throughout for example suppose that you make a solution of sugar in water if you were to take a drop of the sugar solution from anywhere in the solution it would always have the same composition terminology kids learn about solutions and dissolving in chemistry including interesting facts examples solubility saturation concentration and what is a solution solutions play a very important role in many biological laboratory and industrial applications of chemistry of particular importance are solutions involving substances dissolved in water or aqueous solutions a solution in science is a homogenous mixture of two or more substances solutions appear to be one substance but the parts of a solution are not chemically bonded solutions can exist in any phase of matter and the proportions of substances in a solution can vary up to the limit of solubility parts of a solution a solution is a mixture of materials one of which is usually a fluid a fluid is a material that flows such as a liquid or a gas the fluid of a solution is usually the solvent the material other than the solvent is the solute we say that we dissolve the solute into the solvent solutions are mixtures made when a solute dissolves into a solvent learn about solutions in this key stage 3 chemistry guide aged from bbc bitesize introduction a solution is a mixture of two or more substances that stays evenly mixed substances that are combined to form a solution do not change into new substances some examples of solutions include seawater gasoline glass steel and air mixing and separating as they investigate the chemistry of mixtures they learn about hydrophobic hydrophilic and immiscible properties solvents and solutes solubility and saturation phases the tyndall effect and how mixtures are related to real world challenges and solutions solutions are composed of particles the size of an ion or small molecule 0.1 to 200 nm the examples provided above are all considered solutions air is a solution of small gas molecules simple syrup is a solution of sucrose in water and seawater is a solution of ions and water

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solvent is called its solubility

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