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# chapter 9 stoichiometry review flashcards quizlet Apr 28 2024

chapter 9 stoichiometry review flashcards quizlet in the formation of carbon dioxide from carbon monoxide and oxygen how many moles of carbon monoxide are needed to react completely with 7 0 moles of oxygen gas 2 co g  $o_2$  g  $o_2$  g  $o_3$  g  $o_4$  cord to flip 14 mol click the card to flip 1 19 flashcards learn test match

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these quantitative relationships are known as the reaction s stoichiometry a term derived from the greek words stoicheion meaning element and metron meaning measure in this module the use of balanced chemical equations for various stoichiometric applications is explored

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chapter menu resources chapter 9 objectives describe a method for determining which of two reactants is a limiting reactant calculate the amount in moles or mass in grams of a product given the amounts in moles or masses in grams of two reactants one of which is in excess distinguish between theoretical yield actual yield and

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section 1 short answer answer the following questions in the space provided the coefficients in a chemical equation represent the a masses in grams of all reactants and products b relative number of moles of reactants and products c number of atoms of each element in each compound in a reaction

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in this chapter you will learn about quantitative components within a chemical reaction mole ratios within a chemical reaction stoichiometry analysis of a chemical reaction involving moles and mass determine the limiting reactant and the excess reactant in a chemical reaction

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chapter 9 review enhanced introductory college chemistry 9 1 stoichiometry basics and 9 2 mole mass and mass mass calculations write the balanced equation then outline the steps necessary to determine the information requested in each of the following

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objective identify the quantities that are always conserved in chemical reactions stoichiometry is somewhat like a recipe when cooking when baking a recipe is usually used telling the exact amount of each ingredient if you need more you can double or triple the amount thus a recipe is much like a balanced equation stoichiometry is

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we begin this chapter by describing the relationship between the mass of a sample of a substance and its composition we then develop methods for determining the quantities of compounds produced or consumed in chemical reactions and we describe some fundamental types of chemical reactions

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explain the concept of stoichiometry as it pertains to chemical reactions use balanced chemical equations to derive stoichiometric factors relating amounts of reactants and products

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a interpret the equation in terms of numbers of molecules and moles one molecule mole of ethanol reacts with three molecules moles of oxygen gas to yield two molecules moles of carbon dioxide gas and three molecules moles of steam b show that the balanced equation obeys the law of conservation of mass

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